125 g acetone (C_3H_6O)

Temp Change $\Delta T = 450 - 100 = 350 K$

Ketone MW 58.0791 g/mole

1) Moles $= \dfrac{\text{Mass}}{RFM} = \dfrac{125}{}$

$C = (12 \times 3) + 6 + 16 = 58$

$H = (1 \times 6)$

$O = (16)$

moles $= \dfrac{125}{58} = 2.1552$ moles